

**B.Sc. Programme  
(Chemistry Major)**

**Laboratory Manual**

**Volumetric Analysis**



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**TIRUCHIRAPPALLI – 620 002**

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**Weight Calculation :**

Weight of weighing bottle with substance =

Weight of weighing bottle after transferring the substance =

Weight of oxalic acid in 200 ml =

Equivalent weight of oxalic acid =

Normality of oxalic acid =  $\frac{\text{Weight / Litre}}{\text{Equivalent weight}}$ 

$$= \frac{\quad \times 5}{63}$$

Normality of oxalic acid = N

**Titration – I****Standard Oxalic acid Vs Potassium permanganate**

S. No.	Volume of Oxalic acid (ml)	Burette reading (ml)		Volume of KMnO <sub>4</sub> (ml)	Indicator
		Initial	Final		

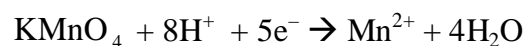
## ESTIMATION OF FERROUS AMMONIUM SULPHATE

### Aim

To estimate the amount of ferrous ammonium sulphate present in the whole of the given solution being supplied with pure crystals of oxalic acid and approximately 0.05N solution of potassium permanganate as the link.

### Principle

Potassium permanganate oxidises  $\text{FeSO}_4$  into  $\text{Fe}_2(\text{SO}_4)_3$  in the presence of dil.  $\text{H}_2\text{SO}_4$



Since, one molecule of ferrous ammonium sulphate loses only one electron. Its equivalent is,

$$\frac{\text{Molecular mass}}{1} = \frac{392}{1} = 392$$

### Procedure :

#### Preparation of standard oxalic acid solution :

Accurately about 0.63g of oxalic acid is weighed in a chemical balance. It is transferred into a 200ml standard flask and dissolved in distilled water. The solution is made upto the mark and shaken well.

Volume of oxalic acid  $V_1 =$

Normality of oxalic acid  $N_1 =$

Volume of potassium permanganate  $V_2 =$

Normality of potassium permanganate  $N_2 =$

$$= \frac{20 \times}{\quad}$$

Normality of potassium permanganate  $N_2 = \quad N$

### Titration – II

#### Made up Ferrous Ammonium Sulphate Vs Standardised potassium permanganate

S. No.	Volume of ferrous ammonium sulphate (ml)	Burette reading (ml)		Volume of $\text{KMnO}_4$ (ml)	Indicator
		Initial	Final		

### **Standardization of Potassium permanganate :**

The given potassium permanganate solution is taken in a clean and rinsed burette. 20ml of standard oxalic acid is pipetted out into a clean conical flask by means of a washed and rinsed pipette. A test tube full of (20ml) dilute sulphuric acid is added and the solution is heated to bearable warmth. The hot solution is titrated against potassium permanganate taken in the burette. The end point is the appearance of pale permanent pink colour. The titrations are repeated to get concordant values. From the titre value, the strength of potassium permanganate solution is calculated.

### **Estimation of Ferrous Ammonium Sulphate**

The given ferrous ammonium sulphate solution is made upto 100ml in a standard flask with distilled water. The flask is shaken well to get uniform concentration. 20ml of this solution is pipetted out into a conical flask by using a washed and rinsed pipette. A test tube full of dilute sulphuric acid is added and the solution is titrated against potassium permanganate taken in the burette. The end point is the appearance of pale permanent pink colour. The titrations are repeated to get concordant values. From the titre value, the strength of ferrous ammonium sulphate and hence its amount is calculated.

Volume of potassium permanganate	$V_1$	=	
Normality of potassium permanganate	$N_1$	=	
Volume of ferrous ammonium sulphate	$V_2$	=	
Normality of ferrous ammonium sulphate	$N_2$	=	
		=	$\frac{x}{20}$
Normality of ferrous ammonium sulphate	$N_2$	=	$N$

**Calculation :**

Equivalent weight of ferrous ammonium sulphate	=	392
Weight / litre of ferrous ammonium sulphate	=	Normality x Equivalent weight
	=	392 x
	=	g
The amount of ferrous ammonium sulphate in 100ml	=	$\frac{\quad}{10}$
	=	g



## Result

The amount of ferrous ammonium sulphate present in the whole of the  
given solution =                      g

**Weight Calculation :**

Weight of weighing bottle with substance =

Weight of weighing bottle after transferring the substance =

Weight of Ferrous ammonium sulphate in 100 ml =

Equivalent weight of Ferrous ammonium sulphate = 392

$$\text{Normality of Ferrous ammonium sulphate} = \frac{\text{Weight / Litre}}{\text{Equivalent weight}}$$
$$= \frac{\quad \times 10}{392}$$

Normality of Ferrous ammonium sulphate = N

**Titration – I****Standard Ferrous Ammonium Sulphate Vs Potassium permanganate**

S. No.	Volume of Ferrous Ammonium Sulphate (ml)	Burette reading (ml)		Volume of KMnO <sub>4</sub> (ml)	Indicator
		Initial	Final		

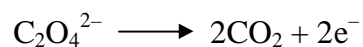
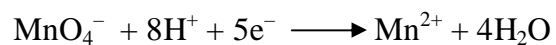
## ESTIMATION OF OXALATE ION

### Aim

To estimate the amount of oxalate ion present in the whole of the given solution being supplied with pure ferrous ammonium sulphate crystals and approximately 0.05N solution of potassium permanganate as the link.

### Principle

Potassium permanganate oxidises oxalate ion according to the following equation



Since an oxalate ion loses two electrons. Its equivalent mass is,

$$\frac{\text{Molecular mass}}{2} = \frac{88}{2} = 44$$

Volume of ferrous ammonium sulphate  $V_1 =$   
 Normality of ferrous ammonium sulphate  $N_1 =$   
 Volume of potassium permanganate  $V_2 =$   
 Normality of potassium permanganate  $N_2 =$   
  
 $N_2 =$   
  
 Normality of potassium permanganate  $N_2 =$  N

### Titration – II

#### Made up oxalic acid Vs Standardised potassium permanganate

S. No.	Volume of ferrous madeup oxalic acid (ml)	Burette reading (ml)		Volume of $\text{KMnO}_4$ (ml)	Indicator
		Initial	Final		

## **Procedure**

### **Preparation of Standard Ferrous Ammonium Sulphate Solution**

Accurately about 1.96g of ferrous ammonium sulphate is weighed and transferred into a 100ml standard flask. It is dissolved in dilute sulphuric acid and the solution is made up to the mark using distilled water. The standard flask is shaken well.

### **Standardization of potassium permanganate**

20ml of standard ferrous ammonium sulphate is pipetted out into a clean conical flask after preliminary washing and rinsing. A test tube full of (20ml) dilute sulphuric acid is added and the solution is titrated against potassium permanganate taken in the burette. The end point is the appearance of pale permanent pink colour. The titrations are repeated to get concordant values. From the titre value, the strength of potassium permanganate solution is calculated.

Volume of $\text{KMnO}_4$	$V_1$	=	
Normality of $\text{KMnO}_4$	$N_1$	=	
Volume of oxalic acid	$V_2$	=	
Normality of oxalic acid	$N_2$	=	
	$N_2$	=	
Normality of oxalic acid	$N_2$	=	N

**Calculation :**

Equivalent weight of oxalate ion	=	44
Weight / litre of oxalate solution	=	Normality x Equivalent weight
	=	x 44
	=	g
Weight of oxalate ion in 100ml of oxalate solution	=	$\frac{\quad}{10}$
	=	g

### **Estimation of Oxalate Ion**

The given Oxalate solution is made upto 100ml in a standard flask with distilled water, 20ml of this solution is pipetted out into a conical flask, added one test tube full of dil.  $\text{H}_2\text{SO}_4$  and heated to bearable warmth. The hot solution is titrated against  $\text{KMnO}_4$  taken in the burette. The end point is the appearance of pale permanent pink colour. The titrations are repeated to get concordant values.

### **Result**

The amount of oxalate ion present in the whole of the given solution =                      g

**Weight Calculation :**

Weight of weighing bottle with substance =

Weight of weighing bottle after transferring the substance =

Weight of Oxalic acid in 200 ml =

Equivalent weight of Oxalic acid = 63

Normality of Oxalic acid =  $\frac{\text{Weight / Litre}}{\text{Equivalent weight}}$ 

$$= \frac{\quad \times 5}{63}$$

Normality of Oxalic acid = N

**Titration – I****Standard Oxalic acid Vs Potassium permanganate**

S. No.	Volume of Oxalic acid (ml)	Burette reading (ml)		Volume of KMnO <sub>4</sub> (ml)	Indicator
		Initial	Final		



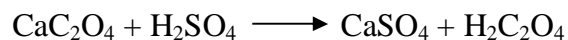
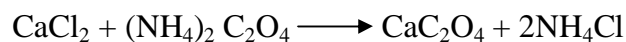
## ESTIMATION OF CALCIUM – DIRECT METHOD

### Aim

To estimate the amount of calcium present in the whole of the given solution, being provided with oxalic acid crystals and approximately 0.05N solution of potassium permanganate as the link.

### Principle

Calcium in calcium chloride is precipitated as calcium oxalate in ammoniacal medium using ammonium oxalate. The precipitate is filtered and dissolved in hot dilute sulphuric acid and the liberated oxalic acid is titrated against potassium permanganate solution.



$$\text{Equivalent mass of calcium} = \frac{\text{Atomic mass}}{2} = 20.04$$

Volume of Oxalic acid  $V_1 =$   
 Normality of Oxalic acid  $N_1 =$   
 Volume of potassium permanganate  $V_2 =$   
 Normality of potassium permanganate  $N_2 =$   
  
 $N_2 =$   
  
 Normality of potassium permanganate  $N_2 =$  N

### Titration – II

**Made up calcium chloride solution Vs standardized potassium permanganate**

S. No.	Volume of calcium chloride solution (ml)	Burette reading (ml)		Volume of $\text{KMnO}_4$ (ml)	Indicator
		Initial	Final		

## **Procedure**

### **Preparation of standard oxalic acid solution**

A standard solution of oxalic acid is prepared by weighing about 0.63g of it accurately, dissolving it in distilled water and making upto 200ml in a standard flask.

### **Standardization of potassium permanganate solution**

The given potassium permanganate solution is taken in a clean and rinsed burette. Exactly 20ml of the standard oxalic acid is pipetted out into a clean conical flask, added equal volume of dilute sulphuric acid and heated to bearable warmth. The hot solution is then titrated against potassium permanganate taken in the burette. The end point is the appearance of pale permanent pink colour. The titrations are repeated to get concordant values.

### **Estimation of Calcium**

The given calcium chloride solution is made upto 100ml in a standard flask with distilled water. 20ml of this solution is pipetted out into a clean 250ml beaker provided with a glass rod and watch glass. A drop by methyl orange is added and the solution is neutralized by the addition of ammonium hydroxide drop by drop. Diluted with water and heated to boiling. About 20ml of ammonium oxalate is taken in a boiling tube and heated to 60-70°C. The hot ammonium oxalate solution is added slowly to the calcium solution with constant stirring. Ammonium hydroxide is then added till the mixture is ammoniacal. The precipitate is digested over a steam bath for half an hour.

Volume of Potassium permanganate	$V_1$	=	
Normality of Potassium permanganate	$N_1$	=	
Volume of calcium solution	$V_2$	=	
Normality of calcium solution	$N_2$	=	
	$N_2$	=	
Normality of calcium solution	$N_2$	=	N

**Calculation :**

Equivalent weight of calcium	=	20.04
Weight / litre of calcium	=	Normality x Equivalent weight
	=	x 20.04
	=	g
Weight of calcium in 100ml of the given solution	=	
	=	g



**Weight Calculation :**

Weight of weighing bottle with substance =

Weight of weighing bottle after transferring the substance =

Weight of Oxalic acid in 200 ml =

Equivalent weight of Oxalic acid = 63

Normality of Oxalic acid =  $\frac{\text{Weight / Litre}}{\text{Equivalent weight}}$ 

$$= \frac{\quad \times 5}{63}$$

Normality of Oxalic acid = N

**Titration – I****Standard Oxalic acid Vs Potassium permanganate**

S. No.	Volume of Oxalic acid (ml)	Burette reading (ml)		Volume of KMnO <sub>4</sub> (ml)	Indicator
		Initial	Final		

## **ESTIMATION OF CALCIUM – INDIRECT METHOD**

### **Aim**

To estimate the amount of calcium present in the whole of the given solution, being provided with pure crystals of oxalic acid and approximately 0.05N solution of potassium permanganate as the link.

### **Principle**

Calcium in calcium chloride solution is precipitated by adding a measured excess of oxalic acid. It is filtered and the unreacted oxalic acid is estimated by titrating with standardized potassium permanganate solution. From the difference in titre values the equivalent of oxalic acid required for calcium solution and hence its weight is calculated.

Volume of Oxalic acid  $V_1 =$   
 Normality of Oxalic acid  $N_1 =$   
 Volume of potassium permanganate  $V_2 =$   
 Normality of potassium permanganate  $N_2 =$   
  
 $N_2 =$   
  
 Normality of potassium permanganate  $N_2 =$  N

### Titration – II

**Made up calcium chloride solution Vs standardised potassium permanganate**

S. No.	Volume of calcium chloride solution (ml)	Burette reading (ml)		Volume of $\text{KMnO}_4$ (ml)	Indicator
		Initial	Final		



## **Procedure**

### **Preparation of standard oxalic acid solution**

A standard solution of oxalic acid is prepared by weighing about 0.63g of it accurately, dissolving it in distilled water and making upto 200ml in a standard flask.

### **Standardization of potassium permanganate solution**

Potassium permanganate is taken in a washed and rinsed burette. Exactly 20ml of oxalic acid is pipetted out into a clean conical flask. Added equal volume of dilute sulphuric acid and heated to bearable warmth. The hot solution is titrated against potassium permanganate taken in the burette. The end point is the appearance of pale permanent pink colour. The titrations are repeated to get concordant values.

$$\begin{aligned}
20\text{ml of } N \text{ oxalic acid} & \equiv \text{ ml of potassium permanganate} \\
40\text{ml of } N \text{ oxalic acid} & \equiv \text{ ml of potassium permanganate} \\
& = \text{ ml of potassium permanganate} \\
\text{Volume of Potassium permanganate} & \equiv \text{ unused oxalic acid} \\
& = \text{ ml} \\
\text{Volume of Potassium permanganate} & \equiv \text{ used oxalic acid} \\
& = \\
& = \text{ ml} \\
\text{ml of } N \text{ of potassium permanganate} & = 20 \text{ ml of calcium chloride solution}
\end{aligned}$$

$$\begin{aligned}
\text{Normality of calcium solution} & = \frac{\quad}{20} \\
& = \quad N
\end{aligned}$$

**Calculation :**

$$\begin{aligned}
\text{Equivalent weight of calcium ion} & = 20.04 \\
\text{Weight / litre of calcium} & = \text{Normality} \times \text{Equivalent weight of calcium ion} \\
& = \quad \times 20.04 \\
& = \quad \text{g} \\
\text{Weight of calcium ion present in 100ml of the given solution} & = \\
& = \quad \text{g}
\end{aligned}$$

## **Estimation of Calcium**

The given calcium chloride solution is made upto 100ml in a standard flask. Exactly 20ml of it is pipetted out into a clean 250ml beaker provided with a glass rod and watch glass. The solution is diluted and heated to boiling. Into the boiling solution exactly 40ml of oxalic acid is pipetted out to precipitate calcium as calcium oxalate. Ammonium hydroxide is then added and the precipitate is digested over a steam bath for half an hour.

A Whatmann No.42 filter paper is placed in a clean funnel. The clear supernatant liquid is transferred into the filter paper and the filtrate is collected in a conical flask. The precipitate is washed many times with distilled water and the washings are also collected in the same conical flask. To the contents of the conical flask, 20ml of dilute sulphuric acid is added and heated to bearable warmth. The hot solution is titrated against potassium permanganate, taken in the burette. The end point is the appearance of pale permanent pink colour. A duplicate is performed simultaneously.

## **Result**

The amount of calcium present in the whole of the given solution

- |     |   |
|-----|---|
| (a) | g |
| (b) | g |

**Weight Calculation :**

Weight of weighing bottle with substance =

Weight of weighing bottle after transferring the substance =

Weight of potassium dichromate in 200 ml =

Equivalent weight of potassium dichromate = 49

Normality of potassium chromate =  $\frac{\text{Weight / Litre}}{\text{Equivalent weight}}$

=  $\frac{\quad \times 5}{49}$

Normality of potassium dichromate = N

**Titration – I****Standard Potassium dichromate Vs sodium thiosulphate solution**

S. No.	Volume of potassium dichromate (ml)	Burette reading (ml)		Volume of thiosulphate (ml)	Indicator
		Initial	Final		

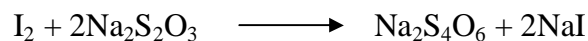
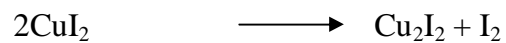
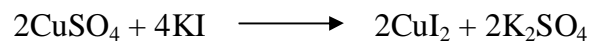
## ESTIMATION OF COPPER

### Aim

To estimate the amount of copper present in the whole of the given solution, being supplied with pure crystals of potassium dichromate and approximately 0.05 N sodium thiosulphate solution.

### Principle

Copper is precipitated as cuprous iodide in acetic acid medium and an equivalent amount of iodine is liberated. The liberated iodine is titrated against sodium thiosulphate.



Equivalent mass of copper = Atomic mass of copper = 63.54

Volume of potassium dichromate  $V_1 =$   
 Normality of potassium dichromate  $N_1 =$   
 Volume of sodium thiosulphate  $V_2 =$   
 Normality of sodium thiosulphate  $N_2 =$   
  
 $=$   
  
 Normality of sodium thiosulphate  $N_2 =$  N

### Titration – II

#### Made up copper sulphate solution Vs standardised sodium thiosulphate

S. No.	Volume of copper sulphate (ml)	Burette reading (ml)		Volume of thiosulphate (ml)	Indicator
		Initial	Final		

## **Procedure**

### **Preparation of standard potassium dichromate solution**

A standard solution of potassium dichromate is prepared by weighing about 0.49g of it accurately, dissolving it in distilled water and making up to 200ml in a standard flask.

### **Standardisation of Sodium Thiosulphate**

The burette is filled with thio solution after washing and rinsing. Exactly 20ml of the standard potassium dichromate solution is pipetted out into a clean conical flask. To this 10ml of dilute hydrochloric acid and 10ml of 10% potassium iodide solution are added and the liberated iodine is titrated against thio sulphate solution. When the solution becomes straw yellow in colour, 1ml of freshly prepared starch solution is added and the titration is continued. The end point is the disappearance of blue colour and appearance of green colour. The titrations are repeated to get concordant values.

Volume of sodium thiosulphate	$V_1$	=	
Normality of sodium thiosulphate	$N_1$	=	
Volume of copper sulphate	$V_2$	=	
Normality of copper sulphate	$N_2$	=	
	$N_2$	=	
Normality of copper sulphate	$N_2$	=	N

**Calculation :**

Equivalent weight of copper	=	63.54
Weight / litre of copper sulphate	=	Normality x Equivalent weight
	=	x 63.54
	=	g
Amount of copper sulphate present in whole of the given solution	=	
	=	g



### **Estimation of Copper**

The given copper sulphate solution is made upto 100ml in a standard flask. Exactly 20ml of this solution is pipetted out into a clean conical flask and ammonium hydroxide is added till a pale blue precipitate is formed. The precipitate is dissolved by adding acetic acid and two drops of acetic acid is added in excess. To this solution 10ml of 10% potassium iodide solution is added and the liberated iodine is titrated against thiosulphate. When the dark brown colour becomes straw yellow, 1ml of starch solution is added and the titration is continued. The end point is the disappearance of blue colour. The titrations are repeated to get concordant values.

### **Result**

The amount of copper present in the whole of the given solution =

**Weight Calculation :**

Weight of weighing bottle with substance =

Weight of weighing bottle after transferring the substance =

Weight of oxalic acid in 200 ml =

Equivalent weight of oxalic acid = 63

Normality of oxalic acid =  $\frac{\text{Weight / Litre}}{\text{Equivalent weight}}$ 

$$= \frac{\quad \times 5}{63}$$

Normality of oxalic acid = N

**Titration – I****Standard Oxalic acid Vs potassium permanganate**

S. No.	Volume of oxalic acid (ml)	Burette reading (ml)		Volume of KMnO <sub>4</sub> (ml)	Indicator
		Initial	Final		

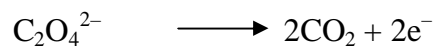
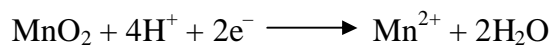
## PERCENTAGE PURITY OF PYROLUSITE

### Aim

To estimate the amount of pure pyrolusite in 100g of the sample using standard oxalic acid and an approximately 0.05N potassium permanganate solution as the link.

### Principle

The pyrolusite ( $\text{MnO}_2$ ) oxidises oxalic acid in the presence of dilute sulphuric acid to carbon dioxide and water



Hence the equivalent mass of  $\text{MnO}_2$

$$= \frac{\text{Molecular mass}}{2} = \frac{86.94}{2} = 43.47$$

Volume of oxalic acid  $V_1 =$   
 Normality of oxalic acid  $N_1 =$   
 Volume of potassium permanganate  $V_2 =$   
 Normality of potassium permanganate  $N_2 =$   
 =  
 Normality of potassium permanganate  $N_2 = N$

### Titration – II

#### Excess Oxalic acid Vs standardised potassium permanganate

S. No.	Volume of oxalic acid (ml)	Burette reading (ml)		Volume of $\text{KMnO}_4$ (ml)	Indicator
		Initial	Final		

## **PROCEDURE**

### **Preparation of standard oxalic acid**

A standard solution of oxalic acid is prepared by weighing about 0.63g of it accurately, dissolving it in distilled water and making upto 200ml in a standard flask.

### **Standardisation of potassium permanganate**

Potassium permanganate solution is taken in a clean and rinsed burette. Exactly 20ml of standard oxalic acid is pipetted out into a clean conical flask and equal volume of dilute sulphuric acid is added. The contents of the conical flask is heated to bearable warmth and titrated against potassium permanganate. The end point is the appearance of pale permanent pink colour. The titrations are repeated to get concordant values.

$$\begin{aligned}
20\text{ml of oxalic acid} & \equiv \text{ml of potassium permanganate} \\
40\text{ml of oxalic acid} & \equiv 2 \times \text{ml of potassium permanganate} \\
& = \text{ml of potassium permanganate} \\
\text{Volume of Potassium permanganate} & \equiv \text{unused oxalic acid} \\
& = \\
& = \text{ml} \\
100 \text{ ml of } N \text{ of potassium permanganate} & = \text{used up oxalic acid} \\
& \text{ [100ml of oxalic acid –} \\
& \text{ 100ml of pyrolusite].} \\
100\text{ml of } N \text{ potassium permanganate} & = 43.47 \text{ g of } N \text{ pyrolusite} \\
\text{ml of } N \text{ potassium permanganate} & = \frac{43.47 \times x}{1000} \\
& = \text{g of pure MnO}_2 \\
0.05\text{g of pyrolusite contain} & \text{g of pure pyrolusite} \\
100 \text{ g of pyrolusite contain} & = \frac{x \times 100}{0.05} \text{ g} \\
& = \%
\end{aligned}$$

### **Determination of percentage purity of Pyrolusite**

About 0.1g of manganese dioxide is weighed accurately and transferred into a clean conical flask. Exactly 40ml of oxalic acid is pipetted out into the conical flask and three test tubes of dilute sulphuric acid ( $\approx$  60ml) is added. The flask is heated gently till all the black particles get dissolved. Care is taken not to allow the mixture to boil, otherwise oxalic acid will volatilise. When the reaction is over, the unreacted oxalic acid is titrated against potassium permanganate taken in the burette. The end point is the appearance of pale permanent pink colour. A duplicate is performed simultaneously.

### **Result**

The percentage purity of pyrolusite =

**Weight Calculation :**

Weight of weighing bottle with substance =

Weight of weighing bottle after transferring the substance =

Weight of oxalic acid in 200 ml =

Equivalent weight of oxalic acid = 63

Normality of oxalic acid =  $\frac{\text{Weight / Litre}}{\text{Equivalent weight}}$ 

$$= \frac{\quad \times 5}{63}$$

Normality of oxalic acid = N

**Titration – I****Standard Oxalic acid Vs Potassium Permanganate**

S. No.	Volume of Oxalic acid (ml)	Burette reading (ml)		Volume of KMnO <sub>4</sub> (ml)	Indicator
		Initial	Final		



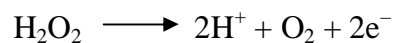
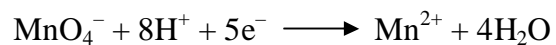
## ESTIMATION OF HYDROGEN PEROXIDE

### Aim

To estimate the volume strength of hydrogen peroxide, being supplied with pure crystals of oxalic acid and an approximately decinormal potassium permanganate solution as the link

### Principle

In acid medium potassium permanganate oxidises hydrogen peroxide as follows



So, the equivalent mass of  $\text{H}_2\text{O}_2$ ,

$$= \frac{\text{Molecular mass}}{2} = \frac{34}{2} = 17$$

Volume of oxalic acid  $V_1 =$   
 Normality of oxalic acid  $N_1 =$   
 Volume of potassium permanganate  $V_2 =$   
 Normality of potassium permanganate  $N_2 =$   
  
 $=$   
  
 Normality of potassium permanganate  $N_2 =$   $N$

### Titration – II

#### Made up Hydrogen peroxide Vs standardised potassium permanganate

S. No.	Volume of hydrogen peroxide (ml)	Burette reading (ml)		Volume of $\text{KMnO}_4$ (ml)	Indicator
		Initial	Final		

## **Procedure**

### **Preparation of standard oxalic acid solution**

A standard solution of oxalic acid is prepared by weighing about 0.63g of it accurately, dissolving it in distilled water and making upto 200ml in a standard flask.

### **Standardisation of potassium permanganate**

Potassium permanganate in solution is taken in a clean and rinsed burette. Exactly 20ml of standard oxalic acid is pipetted out into a clean conical flask and added equal volume of dilute sulphuric acid. The contents of the conical flask is heated to bearable warmth and titrated against potassium permanganate taken in the burette. The end point is the appearance of pale permanent pink colour. The titrations are repeated to get concordant values.

Volume of potassium permanganate  $V_1 =$

Normality of potassium permanganate  $N_1 =$

Volume of hydrogen peroxide  $V_2 =$

Normality of hydrogen peroxide  $N_2 =$

$N_2 =$

Normality of hydrogen peroxide  $N_2 = N$

**Calculation :**

Equivalent weight of hydrogen peroxide  $= 17$

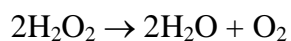
Weight / litre of hydrogen peroxide solution  $= \text{Normality} \times \text{Equivalent weight}$

$= \quad \times 17$

$= \quad \text{g}$

Weight / 200 ml of hydrogen peroxide solution  $= \frac{\quad}{5}$

$= \quad \text{g}$



2 x 34 g of  $\text{H}_2\text{O}_2$  will liberate 32 g of  $\text{O}_2$

$\quad$  g of  $\text{H}_2\text{O}_2$  will liberate  $\frac{32 \times \quad}{68}$

$= \quad \text{g}$

32 g of oxygen at NTP will occupy 22400 ml

$\quad$  g of oxygen at NTP will occupy  $\frac{22400 \times \quad}{32}$

$= \quad \text{ml}$

### **Estimation of Hydrogen Peroxide**

The given hydrogen peroxide is made upto 100ml in a standard flask with distilled water. Exactly 20ml of this solution is pipetted out into a clean conical flask. About 20ml of dilute sulphuric acid is added to it and titrated against standardized potassium permanganate taken in the burette. The end point is the appearance of pale permanent pink colour. The titrations are repeated to get concordant values.

### **Result**

Volume strength of hydrogen peroxide =            ml

**Weight Calculation :**

Weight of weighing bottle with substance =

Weight of weighing bottle after transferring the substance =

Weight of oxalic acid in 200 ml =

Equivalent weight of oxalic acid = 63

Normality of oxalic acid =  $\frac{\text{Weight / Litre}}{\text{Equivalent weight}}$ 

$$= \frac{\quad \times 5}{63}$$

Normality of oxalic acid = N

**Titration – I****Standard Oxalic acid Vs potassium permanganate**

S. No.	Volume of Oxalic acid (ml)	Burette reading (ml)		Volume of KMnO <sub>4</sub> (ml)	Indicator
		Initial	Final		

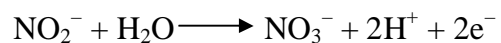
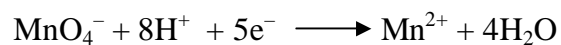
## ESTIMATION OF NITRITE ION

### Aim

To estimate the amount of nitrite ion present in the whole of the given solution using standard oxalic acid and approximately 0.05N potassium permanganate solution as the link.

### Principle

Nitrite is oxidized to nitrate by a warm acidified solution of potassium permanganate



Hence, the equivalent mass of nitrite,

$$= \frac{\text{Molecular mass}}{2} = \frac{46}{2} = 23$$

Volume of oxalic acid  $V_1 =$   
 Normality of oxalic acid  $N_1 =$   
 Volume of potassium permanganate  $V_2 =$   
 Normality of potassium permanganate  $N_2 =$   
  
 $=$   
 Normality of potassium permanganate  $N_2 = N$

### Titration – II

#### Standardised potassium permanganate Vs made up nitrite solution

S. No.	Volume of $\text{KMnO}_4$ (ml)	Burette reading (ml)		Volume of sodium nitrite solution (ml)	Indicator
		Initial	Final		



## **Procedure**

### **Preparation of standard oxalic acid solution**

About 0.63g of oxalic acid is weighed accurately, transferred into a 200ml standard flask and dissolved in distilled water. The solution is made upto 200ml.

### **Standardisation of potassium permanganate solution**

Exactly 20ml of oxalic acid is pipetted out into a clean conical flask. Equal volume of dilute sulphuric acid is added and heated to bearable warmth. The hot solution is titrated against potassium permanganate taken in the burette. The end point is the appearance of pale permanent pink colour. The titrations are repeated to get concordant values.

Volume of potassium permanganate	$V_1$	=	
Normality of potassium permanganate	$N_1$	=	
Volume of sodium nitrite solution	$V_2$	=	
Normality of sodium nitrite solution	$N_2$	=	
		=	
Normality of sodium nitrite solution	$N_2$	=	N

**Calculation :**

Equivalent weight of nitrite ion	=	23
Weight / litre of nitrite solution	=	Normality x Equivalent weight
	=	x 23
	=	g
Weight of nitrite ion present in the whole of the given solution	=	g

### **Estimation of Nitrite Ion**

The given nitrite solution is made upto 100ml in a standard flask. The made up nitrite solution is taken in a clean and rinsed burette. Exactly 20ml of potassium permanganate solution is pipetted out into a clean beaker. Equal volume of dilute sulphuric acid is added and heated to 40°C after diluting the solution to 100ml. The tip of the burette is allowed to remain under the surface of potassium permanganate and the nitrite solution is added drop by drop. The solution is stirred using a glass rod after addition of each drop of nitrite solution. The end point is the disappearance of pink colour. The titrations are repeated to get concordant values.

### **Result**

The amount of nitrite present in the whole of the given solution =                      g

**Weight Calculation :**

Weight of weighing bottle with substance =

Weight of weighing bottle after transferring the substance =

Weight of potassium chloride in 100 ml =

Equivalent weight of potassium chloride = 74.46

Normality of potassium chloride =  $\frac{\text{Weight / Litre}}{\text{Equivalent weight}}$ 

$$= \frac{\quad \times 10}{74.46}$$

Normality of potassium chloride = N

**Titration – I****Standard Potassium chloride Vs silver nitrate**

S. No.	Volume of potassium chloride (ml)	Burette reading (ml)		Volume of silver nitrate (ml)	Indicator
		Initial	Final		

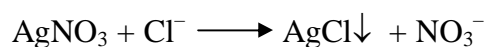
## ESTIMATION OF CHLORIDE ION – MOHR'S METHOD

### Aim

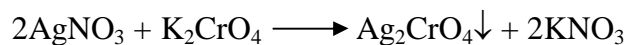
To estimate the amount of chloride present in the whole of the given solution, being supplied with an approximately 0.05N silver nitrate solution and pure crystals of potassium chloride

### Principle

It is an example of precipitation reaction



The completion of the reaction is observed by using potassium chromate as the indicator. The yellow colour of potassium chromate changes into reddish brown due to the formation of silver chromate.



Volume of potassium chloride  $V_1 =$   
 Normality of potassium chloride  $N_1 =$   
 Volume of silver nitrate solution  $V_2 =$   
 Normality of silver nitrate solution  $N_2 =$   
  
 $=$   
  
 Normality of silver nitrate solution  $N_2 = N$

**Titration – II**

**Made up chloride solution Vs standardised silver nitrate**

S. No.	Volume of chloride solution (ml)	Burette reading (ml)		Volume of silver nitrate (ml)	Indicator
		Initial	Final		

## **Procedure**

### **Preparation of standard potassium chloride solution**

About 0.375 g of potassium chloride is accurately weighed and it is transferred into a 100ml standard flask. The crystals are dissolved in distilled water and made upto the mark.

### **Standardisation of Silver Nitrate Solution**

Exactly 20ml of potassium chloride solution is pipetted out into a clean conical flask and 2ml of 5% potassium chromate solution is added as indicator. This solution is titrated against silver nitrate solution taken in the burette. During each addition of silver nitrate solution, the conical flask is shaken well. Nearing the end point, the silver chloride gets coagulated at the bottom of the conical flask. The addition of silver nitrate is done drop by drop until the supernatant solution gets reddish brown tinge. The titrations are repeated to get concordant values.

Volume of silver nitrate solution	$V_1$	=	
Normality of silver nitrate solution	$N_1$	=	
Volume of chloride solution	$V_2$	=	
Normality of chloride solution	$N_2$	=	
	$N_2$	=	
Normality of chloride solution	$N_2$	=	N

**Calculation :**

Equivalent weight of chloride	=	35.46
Weight / litre of chloride	=	Normality x Equivalent weight
	=	x 35.46
	=	g
Weight of chloride present in the whole of the given solution	=	$\frac{\quad}{10}$
	=	g





**Weight Calculation :**

Weight of weighing bottle with substance =

Weight of weighing bottle after transferring the substance =

Weight of potassium chloride in 100 ml =

Equivalent weight of potassium chloride = 74.46

Normality of potassium chloride =  $\frac{\text{Weight / Litre}}{\text{Equivalent weight}}$

=  $\frac{\quad \times 10}{74.46}$

Normality of potassium chloride = N

**Titration – I****Standard Potassium chloride Vs silver nitrate**

S. No.	Volume of potassium chloride (ml)	Burette reading (ml)		Volume of silver nitrate (ml)	Indicator
		Initial	Final		

Volume of potassium chloride  $V_1 =$

Normality of potassium chloride  $N_1 =$

Volume of silver nitrate  $V_2 =$

Normality of silver nitrate  $N_2 =$

=

Normality of silver nitrate  $N_2 =$  N

## ESTIMATION OF COMMERCIAL HYDROCHLORIC ACID

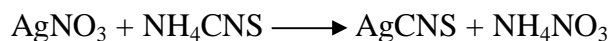
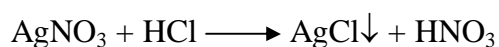
### - VOLHARD'S METHOD

#### **Aim**

To determine the normality of commercial hydrochloric acid by Volhard's method. Being supplied with 0.05N solution of silver nitrate ammonium, thiocyanate and pure crystals of potassium chloride.

#### **Principle**

Volhard's method is applicable to the estimation of chloride in neutral (or) acid medium. In this method the known volume of chloride is treated with known excess of silver nitrate solution. The unreacted silver nitrate is treated with standard ammonium thiocyanate solution. From the titre value, the volume of thiocyanate equivalent to the silver nitrate that has reacted with chloride solution is calculated from which the strength of chloride solution and hence the weight of chloride in the whole of the given solution is calculated.



The indicator used is ferric alum solution. At the end point a reddish brown colour develops in the solution.

**Titration – II****Standardised silver nitrate Vs Ammonium thiocyanate**

S. No.	Volume of silver nitrate (ml)	Burette reading (ml)		Volume of Ammonium thiocyanate (ml)	Indicator
		Initial	Final		

Volume of silver nitrate  $V_1 =$

Normality of silver nitrate  $N_1 =$

Volume of Ammonium thiocyanate  $V_2 =$

Normality of Ammonium thiocyanate  $N_2 =$

$=$

Normality of Ammonium thiocyanate  $N_2 = N$

20 ml N silver nitrate  $\equiv$  ml of Ammonium thiocyanate

40 ml N silver nitrate  $\equiv$  x 2 of Ammonium thiocyanate

$=$  ml

**Titration – III****Made up hydrochloric acid Vs standardised ammonium thiocyanate**

S. No.	Volume of hydrochloric acid (ml)	Burette reading (ml)		Volume of ammonium thiocyanate (ml)	Indicator
		Initial	Final		

## **Procedure**

### **Standardisation of Silver Nitrate Solution**

Exactly 20ml of potassium chloride solution is pipetted out into a clean conical flask and 2ml of 5% potassium chromate solution is added as indicator. This solution is titrated against silver nitrate solution taken in the burette. During each addition of silver nitrate solution, the conical flask is shaken well. Nearing the end point, the silver chloride gets coagulated at the bottom of the conical flask. The addition of silver nitrate is done drop by drop until the supernatant solution gets reddish brown tinge. The titrations are repeated to get concordant values.

### **Standardization of Ammonium Thiocyanate**

Exactly 20ml of the standardised silver nitrate solution is pipetted out into a clean conical flask and 5ml of dilute nitric acid and one ml ferric alum indicator are added. The solution is then titrated against ammonium thiocyanate solution taken in the burette. The end point is the appearance of reddish tinge. The titrations are repeated to get concordant values.

### **Estimation of Hydrochloric Acid**

One ml of commercial hydrochloric acid is made upto 200ml in a standard flask with distilled water. 20ml of this solution is pipetted out into a clean beaker provided with a glass rod and watch glass. About 5ml of dilute nitric acid is added and 40ml of silver nitrate solution is pipetted out into the same beaker. The mixture is stirred well and the precipitated silver chloride is filtered through Whatmann No.41 filter paper. The precipitate is washed well with cold water. The filtrate and washing are collected in a conical flask. About one ml of ferric alum is added to the filtrate, and it is titrated against standardised ammonium thiocyanate solution taken in the burette. The end point is the appearance of reddish tinge. A duplicate is performed side by side.

$$\begin{aligned}
 \text{Volume of Ammonium thiocyanate} & \equiv \text{ unused silver nitrate} \\
 & = \text{ ml} \\
 \text{Volume of Ammonium thiocyanate} & \equiv \text{ used up silver nitrate} \\
 & = \\
 & = \text{ ml} \\
 \text{used up silver nitrate} & = \frac{\text{ml of Ammonium thiocyanate}}{20\text{ml of made up hydrochloric acid}}
 \end{aligned}$$

$$\begin{aligned}
 \text{Normality of chloride solution} & = \frac{\quad}{20} \\
 & = \text{ N}
 \end{aligned}$$

$$\text{Equivalent weight of chloride} = 35.46$$

$$\begin{aligned}
 \text{Weight / litre of chloride solution} & = \text{Normality} \times \text{Equivalent weight} \\
 & = \quad \times 35.46 \\
 & = \text{ g}
 \end{aligned}$$

$$\begin{aligned}
 \text{Weight of chloride present in 200ml} \\
 \text{of chloride solution} & = \frac{\quad}{5} \\
 & = \text{ g}
 \end{aligned}$$

\_\_\_\_\_ g of chloride present in 1ml of commercial hydrochloric acid

1 ml of commercial hydrochloric acid contains \_\_\_\_\_ g of chloride ion.

$$\begin{aligned}
 \left. \begin{array}{l} \text{Weight of chloride ion present in 1000ml} \\ \text{of commercial hydrochloric acid} \end{array} \right\} & = \quad \times 1000 \\
 & = \text{ g}
 \end{aligned}$$

$$\begin{aligned}
 \text{Normality of commercial hydrochloric acid} & = \frac{\text{Weight / Litre}}{\text{Equivalent weight}} \\
 & = \frac{\quad}{35.46} \\
 & = \text{ N}
 \end{aligned}$$

**Result**

The normality of the commercial hydrochloric acid is = \_\_\_\_\_ N

**Weight Calculation :**

Weight of weighing bottle with substance =

Weight of weighing bottle after transferring the substance =

Weight of zinc sulphate in 100ml =

Equivalent weight of zinc sulphate = 143.77

Normality of zinc sulphate =  $\frac{\text{Weight / Litre}}{\text{Equivalent weight}}$

=  $\frac{\quad \times 10}{143.77}$

Normality of zinc sulphate = N

**Titration – I****Standard zinc sulphate Vs potassium ferrocyanide solution**

S. No.	Volume of zinc sulphate solution (ml)	Burette reading (ml)		Volume of potassium ferrocyanide solution (ml)	Indicator
		Initial	Final		



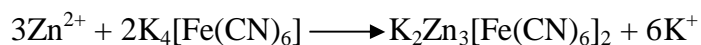
## ESTIMATION OF ZINC

### Aim

To estimate the amount of zinc present in the whole of the given solution, being supplied with approximately 0.05N potassium ferrocyanide solution and pure zinc sulphate crystals.

### Principle

Zinc ions and ferrocyanide ions react in neutral or acid medium as follows.



The zinc ions get precipitated as potassium zinc ferrocyanide. In order to make the reaction as oxidation – reduction a small quantity of potassium ferricyanide is added to the ferrocyanide solution. Equivalent mass of zinc = atomic mass of zinc = 65.38

Volume of zinc sulphate  $V_1 =$   
 Normality of zinc sulphate  $N_1 =$   
 Volume of potassium ferrocyanide  $V_2 =$   
 Normality of potassium ferrocyanide  $N_2 =$   
  
 $=$   
  
 Normality of potassium ferrocyanide  $N_2 =$  N

### Titration – II

#### Made up zinc sulphate solution Vs standardised potassium ferrocyanide

S. No.	Volume of zinc sulphate solution (ml)	Burette reading (ml)		Volume of potassium ferrocyanide solution (ml)	Indicator
		Initial	Final		

## **Procedure**

### **Preparation of standard zinc sulphate solution**

A decinormal solution of zinc sulphate is prepared by weighing about 0.725g of zinc sulphate accurately, dissolving it in distilled water and making upto 100ml in a standard flask.

### **Standardization of Ferrocyanide Solution**

Exactly 20ml of zinc sulphate solution is pipetted out into a clean conical flask. To that solution about 20ml of dilute sulphuric acid, one gram of ammonium sulphate and 2 to 3 drops of diphenylamine indicator are added. The solution is titrated against potassium ferrocyanide solution taken in the burette. The end point is the colour change from blue to green. The titrations are repeated to get concordant values.

Volume of potassium ferrocyanide  $V_1 =$

Normality of potassium ferrocyanide  $N_1 =$

Volume of made up zinc sulphate  $V_2 =$

Normality of made up zinc sulphate  $N_2 =$

$=$

Normality of made up zinc sulphate  $N_2 = N$

**Calculation :**

Equivalent weight of zinc  $= 65.38$

Weight / litre of zinc  $= \text{Normality} \times \text{Equivalent weight}$

$= \quad \times 65.38$

$= \quad \text{g}$

Weight of zinc present in the whole of the given solution  $=$

$= \quad \text{g}$



## Titration – I

### Hard water Vs standardised EDTA solution

S. No.	Volume of hard water (ml)	Burette reading (ml)		Volume of EDTA solution (ml)	Indicator
		Initial	Final		

Strength of EDTA =

1000ml of 1M EDTA  $\equiv$  40.08 g of  $\text{Ca}^{2+}$  ions

ml of 0.01M EDTA =  $\frac{40.08 \times 0.01 \times}{1000 \times 1}$

ml of 0.01M EDTA =  $\times 10^{-3}$  g of  $\text{Ca}^{2+}$

60ml of water contains g of  $\text{Ca}^{2+}$  ions.

$10^6$ ml of water contains  $\frac{\times 10^{-3} \times 10^6}{60}$  g of  $\text{Ca}^{2+}$  ions.

= g of  $\text{Ca}^{2+}$  ions.

40.08g of  $\text{Ca}^{2+}$  ions present in 100.08g of  $\text{CaCO}_3$

$\therefore$  g of  $\text{Ca}^{2+}$  ions present in  $\frac{\times 100.08}{40.08}$  g of  $\text{CaCO}_3$

= g of  $\text{CaCO}_3$

= PPM



**Weight Calculation :**

Weight of weighing bottle with substance =

Weight of weighing bottle after transferring the substance =

Weight of potassium dichromate in 200 ml =

Equivalent weight of potassium dichromate = 49

Normality of potassium dichromate =  $\frac{\text{Weight / Litre}}{\text{Equivalent weight}}$

=  $\frac{\quad \times 5}{49}$

Normality of potassium dichromate = N

**Titration – I****Ferrous Ferric mixture Vs Standard Potassium dichromate**

S. No.	Volume of ferrous ferric mixture (ml)	Burette reading (ml)		Volume of standard $\text{K}_2\text{Cr}_2\text{O}_7$ (ml)	Indicator
		Initial	Final		



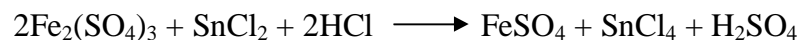
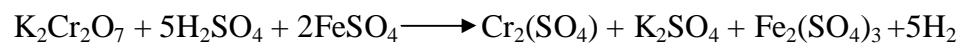
## ESTIMATION OF FERROUS AND FERRIC IRON IN A MIXTURE

### Aim

To estimate the amount of ferrous and ferric iron in a mixture, being supplied with pure crystals of potassium dichromate.

### Principle

The amount of ferrous iron in a mixture of ferrous and ferric iron is determined by direct titration with standard potassium dichromate solution. The ferric iron is reduced to ferrous iron by means of stannous chloride and the total ferrous iron is estimated using potassium dichromate.



Volume of potassium dichromate  $V_1 =$   
 Normality of potassium dichromate  $N_1 =$   
 Volume of ferrous ferric mixture  $V_2 =$   
 Normality of ferrous ferric mixture  $N_2 =$   
 =  
 Normality of ferrous ferric mixture  $N_2 = N$   
 Equivalent weight of ferrous iron  $= 55.85$   
 Weight / Litre of ferrous iron  $= \quad \times 55.85$   
 =  $g$

### Titration – II

#### Ferrous – Ferric mixture after reduction Vs standardised potassium dichromate

S. No.	Volume of ferrous ferric mixture (ml)	Burette reading (ml)		Volume of standard $K_2Cr_2O_7$ (ml)	Indicator
		Initial	Final		

## **Procedure**

### **Preparation of Standard Potassium Dichromate Solution**

Accurately about 0.5g of potassium dichromate, is weighed in a chemical balance, transferred to a 200ml standard flask. It is dissolved in distilled water and made upto the mark.

### **Estimation of Ferrous Iron**

Exactly 20ml of the given ferrous-ferric mixture is pipetted out into a clean conical flask. Added 20ml of a mixture of sulphuric acid and phosphoric acid and three drops of diphenyl amine indicator and titrated versus standard potassium dichromate solution taken in the burette. The end point is the appearance of bluish violet colour. The titrations are repeated to get concordant values.

Volume of potassium dichromate  $V_1 =$

Normality of potassium dichromate  $N_1 =$

Volume of ferrous-ferric mixture  $V_2 =$

Normality of ferrous-ferric mixture  $N_2 =$

$=$

Normality of ferrous-ferric mixture  $N_2 = N$

**Calculation :**

Equivalent weight of ferrous iron  $= 55.85$

Weight / litre of ferrous iron  $= \text{Normality} \times \text{Equivalent weight}$

$= \quad \times 55.85$

$= \quad \text{g}$

Amount of ferric iron present in the whole of the given solution  $= (\text{Amount of total ferrous iron}) - (\text{Amount of ferrous ion})$

$=$

Amount of ferric iron present in whole of the given solution  $= \quad \text{g}$

### **Reduction of Ferric Iron to ferrous Iron**

Exactly 20ml of the given ferrous – ferric mixture is pipetted out into a clean conical flask and 5ml of concentrated hydrochloric acid is added to it. The solution is heated to boiling, and to the hot solution freshly prepared stannous chloride is added till the yellow colour of the solution disappears. Two drops of stannous chloride is added in excess. The solution is diluted to about 150ml, cooled and then 10ml of a saturated solution of mercuric chloride is added to it. A silky white precipitate is formed (If a black precipitate is formed or no precipitate, the reduction is incomplete and the solution is discarded).

The reduced solution containing ferrous iron is mixed with 20ml of mixture of sulphuric acid and phosphoric acid and three drops of diphenyl amine indicator, and titrated against standard potassium dichromate as described above.

### **Result**

1. The amount of ferrous iron present in the whole of the given solution = g
2. The amount of ferric iron present in the whole of the given solution = g

**Weight Calculation :**

Weight of arsenious oxide in 100 ml =

Equivalent weight of arsenious oxide = 49.45

Normality of arsenious oxide =  $\frac{\text{Weight / Litre}}{\text{Equivalent weight}}$

=  $\frac{\quad \times 10}{49.45}$

Normality of arsenious oxide = N

**Titration – I****Standard arsenious oxide Vs Iodine solution**

S. No.	Volume of arsenious oxide (ml)	Burette reading (ml)		Volume of iodine solution (ml)	Indicator
		Initial	Final		

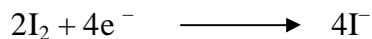
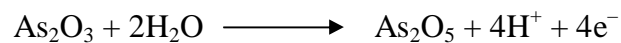
## ESTIMATION OF ARSENIOS OXIDE

### Aim

To estimate the amount of arsenious oxide present in the whole of the given solution, being supplied with pure arsenious oxide and approximately 0.05N iodine solution.

### Principle

Iodine oxidises arsenious oxide. Since the reaction is reversible, the hydrogen iodide formed must be removed. Sodium bicarbonate is added in excess which removes hydrogen iodide and it has no action on iodine. As arsenious oxide is not freely soluble in water, it is dissolved in sodium hydroxide to form sodium arsenite.



$$\text{Equivalent mass of As}_2\text{O}_3 = \frac{\text{Molecular mass}}{4} = \frac{197.82}{4} = 49.45$$

Volume of arsenious oxide solution  $V_1 =$

Normality of arsenious oxide solution  $N_1 =$

Volume of iodine solution  $V_2 =$

Normality of iodine solution  $N_2 =$

$=$

Normality of iodine solution  $N_2 = N$

### Titration – II

#### Made up Arsenious oxide solution Vs standardised Iodine solution

S. No.	Volume of arsenious oxide (ml)	Burette reading (ml)		Volume of iodine solution (ml)	Indicator
		Initial	Final		



## **Procedure**

### **Preparation of Standard Arsenious Oxide Solution**

Accurately about 0.25g of arsenious oxide is weighed in a beaker and a few drops of distilled water is added to make it into a paste. Then sodium hydroxide pellets are added one by one and stirred well to dissolve arsenious oxide. The clear solution is transferred into a 100ml standard flask and the beaker is washed repeatedly with distilled water and the washings are transferred into the flask. The solution is made upto the mark with the distilled water.

### **Standardization of Iodine Solution**

The given iodine solution is taken in a clean and rinsed burette. The standard arsenious oxide solution is filled in another burette and exactly 20ml of it is buretted out into a clean conical flask. A drop of phenolphthalein is added and then the solution is neutralized with dilute hydrochloric acid and 3g of sodium bicarbonate is added, shaken well to dissolve the solid and then added 1ml of starch. The solution is titrated against iodine solution and the end point is the appearance of blue colour. The titrations are repeated to get concordant values.

Volume of iodine solution	$V_1$	=	
Normality of iodine solution	$N_1$	=	
Volume of made up arsenious oxide	$V_2$	=	
Normality of made up arsenious oxide	$N_2$	=	
	$N_2$	=	
Normality of made up arsenious oxide	$N_2$	=	N

**Calculation :**

Equivalent weight arsenious oxide	=	49.45
Weight / litre of arsenious oxide	=	Normality x Equivalent weight
	=	x 49.45
	=	g
Amount of arsenious oxide in 100 ml	=	$\frac{\quad}{10}$
	=	g

